

(19)



Europäisches Patentamt
European Patent Office
Office européen des brevets

(11)



EP 0 565 364 B1

(12)

EUROPEAN PATENT SPECIFICATION

(45) Date of publication and mention
of the grant of the patent:
11.11.1998 Bulletin 1998/46

(51) Int Cl. 6: C09B 39/00

(21) Application number: 93302737.7

(22) Date of filing: 07.04.1993

(54) Aluminosilicates

Aluminosilikate

Aluminosilicates

(84) Designated Contracting States:
AT BE CH DE DK ES FR GB GR IE IT LI NL PT SE

(30) Priority: 09.04.1992 EP 92303147

(43) Date of publication of application:
13.10.1993 Bulletin 1993/41

(73) Proprietor: CROSFIELD LIMITED
Warrington, Cheshire WA5 1AB (GB)

(72) Inventor: Araya, Abraham,
Unilever Res. Port Sunlight Lab.
Wirral, Merseyside L63 3JW (GB)

(74) Representative:
Collingwood, Anthony Robert et al
ICI Group Intellectual Property
P.O. Box 11
The Heath
Runcorn Cheshire WA7 4QE (GB)

(56) References cited:
EP-A- 0 384 070 WO-A-94/26662
DE-A- 2 026 401

EP 0 565 364 B1

Note: Within nine months from the publication of the mention of the grant of the European patent, any person may give notice to the European Patent Office of opposition to the European patent granted. Notice of opposition shall be filed in a written reasoned statement. It shall not be deemed to have been filed until the opposition fee has been paid. (Art. 99(1) European Patent Convention).

DescriptionField of the invention

5 The invention describes the preparation of alkali metal aluminosilicates, in particular crystalline aluminosilicates having the P-structure. These materials are of value as a component in detergent formulations in which they remove calcium and magnesium hardness ions by ion exchange. They also have other properties which provide additional benefits when used in detergent formulations. These aluminosilicates will be referred to as zeolite P in this description.

Background of the invention

10 While the utility of zeolite P in detergent formulations has been acknowledged, for example in European Patent Application 0384070 (Unilever), they must be manufactured by a commercially effective process in order to become available on the market place. Thus, while the properties of a material may make it a useful commercial commodity, 15 its manufacture has to be optimised for large scale production.

15 The zeolite P class includes a series of synthetic zeolite phases which may be in cubic configuration (also termed B or P_c) or tetragonal configuration (also termed P_1) but is not limited to these forms. The structure and characteristics of the zeolite P class are given in "Zeolite Molecular Sieves" of Donald W Breck (published 1974 and 1984 by Robert E Kriger of Florida USA). The zeolite P class has the typical oxide formula:

20



25 The present invention provides a process for preparing crystalline P-zeolites having Si:Al ratios from 0.9 to 1.33, which are particularly effective as detergent builders, preferably ratios of 1.15:1 and below and more preferably 1.07: 1 and below.

30 M is an n-valent cation which for this Invention is an alkali-metal, that is lithium, potassium, sodium, caesium or rubidium with sodium and potassium being preferred and sodium being the cation normally used in commercial processes.

35 Thus sodium may be present as the major cation with another alkali metal present in a minor proportion to provide a specific benefit.

35 The process of the present invention allows the production of P-zeolite with the above defined ratios at economic yield. During the formation of the crystalline zeolite the reacted medium passes through a gel stage. A preferred embodiment of the invention is the production of P-zeolites at a silicon:aluminium molar ratio of 1.00 ± 0.05 from a synthesis gel with the same compositional ratio.

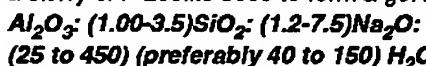
40 This feature results in the desired product and a filtrate liquor containing essentially only sodium hydroxide. This liquor can be recirculated, after optional concentration, e.g. by evaporation, to the preparation of the initial reactant solutions. Another advantage of producing P-zeolites from a synthesis gel with the same silicon:aluminium ratio is that all the sodium silicate and the sodium aluminate are used thus optimising the consumption of the starting material.

45 Moreover, for detergents applications there is a need for a product with high oil absorption.

General description of the invention

45 It is a first object of the present Invention to provide a process for the preparation of P zeolite having the oxide formula $M_n^{2+}O \cdot Al_2O_3 \cdot (1.80-2.66)SiO_2 \cdot yH_2O$ wherein y is the water content,

50 i. a sodium aluminate solution at a temperature of at least 25°C is mixed with a sodium silicate solution and a temperature of at least 25°C in an adequately vigorously stirred vessel in the presence of an effective amount of a slurry of P zeolite seed to form a gel having the composition,



55 ii. ageing the gel at a temperature above about 25°C with adequate vigorous stirring to maintain solids in suspension for a period of at least about 0.1 hours, and

iii. separating the P zeolite product, washing and drying.

The product is washed in order to remove excess electrolytes such as ca.

The sources of the seed zeolite P is not critical, although preferably it is added to the reactants as a previously prepared slurry. Alternatively a crystallised slurry from a previous reaction may be used. Additionally the ratio of the P-zeolite seed is not critical and a P-zeolite having a Si:Al ratio above 1.33 can be used.

That is to say the seed crystalline form drives the reaction to the P-form and the product ratio is attained from the gel ratio of the reactants. The processes of the invention can be operated at the scale required for commercial activities.

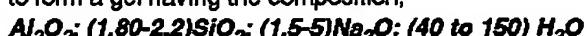
The ratios and concentrations of the aluminates and silicates used are not critical, but must be known to allow calculation of the gel and product compositions. The temperatures used will usually be well above the minimum temperature of 25°C and are preferably above about 60°C. The gel composition will preferably contain about 40 to 150 moles of water per mole alumina.

The range of the silicon:aluminium ratio of the gel is wider than the corresponding ratio of the product since the silicon:aluminium ratio of the product is always between 2 and the corresponding ratio of the gel.

The period for ageing the gel to form the product zeolite is dependent on composition and temperature and will preferably be complete in about 10 hours, although the process will more preferably be completed within 5 hours for a commercial process.

Since it is particularly advantageous to produce P-zeolites at a silicon:aluminilum molar ratio of 1.00±0.05 from a synthesis gel with the same compositional ratio, it is a second object of the present invention to provide a process for the preparation of P zeolite having the oxide formula $M_n^2O \cdot Al_2O_3 \cdot (1.90-2.10)SiO_2 \cdot yH_2O$ wherein y is the water content, comprising the steps of:

I. a sodium aluminate solution having a temperature of at least 60°C is mixed with a sodium silicate solution at a temperature of at least 60°C in a stirred vessel in the presence of an effective amount of a slurry of P zeolite seed to form a gel having the composition,



II. ageing the gel at a temperature above about 60°C with stirring to maintain solids in suspension for a period of least about 0.1 hour, and

iii. separating the P zeolite product, washing and drying.

The process described is a particularly favourable route for preparing P zeolites having improved properties for use as detergent builders. These P zeolites have effective binding capacities for calcium of above 150, preferably above 160 mg/g, good rate of calcium uptake i.e. below 25 seconds, preferably below 10 secs and effective oil absorption i.e. at least 50g/100g zeolite and in the range to 100g/100g zeolite. Their d_{50} was in the range of about 0.85 to about 2.00 micron and was therefore particularly suitable as detergent builders.

Standard procedures:

In the characterisation of the zeolite-P materials prepared by the present process the following methods were used.

i. Particle size: The average particle size (microns) was measured by a Malvern Mastersizer (Trade Mark) obtainable from Malvern Instruments, England and expressed as the d_{50} , i.e. 50% by weight of the particles have a diameter smaller than the diameter quoted. The definitions d_{50} and d_{90} may also be used in association with the appropriate figure. Prior to measuring the particle size, the sample is dried as a thin layer (about 1.5 mm) at 110°C for two hours in a fan oven.

ii. Calcium uptake rate (CUR): The rate of removal of Ca^{++} ions from a wash liquor is an important characteristic of a detergency builder. The time, in seconds, is determined for a zeolite, at a concentration of 1.48gdm⁻³ and a temperature of 25°C, to reduce the calcium ion concentration in a 0.01M sodium chloride solution from an initial value of 2×10^{-3} M to 10^{-5} M. The zeolite was first equilibrated to constant weight over saturated sodium chloride solution and the water content measured.

iii. Calcium effective binding capacity (CEBC): The CEBC was measured in the presence of a background electrolyte to provide a realistic indicator of calcium ion uptake in a wash liquor environment. A sample of each zeolite was first equilibrated to constant weight over saturated sodium chloride solution and the water content measured. Each equilibrated sample was dispersed in water (1cm³) in an amount corresponding to 1 g dm⁻³ (dry), and the resulting dispersion (1 cm³) was injected into a stirred solution, consisting of 0.01M NaCl solution (50 cm³) and 0.05M $CaCl_2$ (3.923 cm³), therefore producing a solution of total volume 54.923 cm³. This corresponded to a

5 concentration of 200 mg CaO per litre, i.e. just greater than the theoretical maximum amount (197 mg) that can be taken up by a zeolite of Si:Al ratio 1.00. The change in Ca²⁺ ion concentration was measured by using a Ca²⁺ ion selective electrode, the final reading being taken after 15 minutes. The temperature was maintained at 25°C throughout. The Ca²⁺ ion concentration measured was subtracted from the initial concentration, to give the effective calcium binding capacity of the zeolite sample as mg CaO/g zeolite.

10 iv. Oil absorption: This was determined by the ASTM spatula rub-out method (American of Test Material Standards D281). The test is based on the principle of mixing linseed oil with the particulate zeolite by rubbing with a spatula on a smooth surface until a stiff putty like paste is formed which will not break or separate when it is cut with the spatula. The weight of oil used is then put into the equation:

$$15 \text{ Oil absorption (OA)} = \frac{g \text{ oil absorbed} \times 100}{\text{wt. of zeolite (gms)}} \\ = g \text{ oil}/100g \text{ zeolite}$$

20 In order to measure the oil absorption of the product, the washed filter cakes obtained after crystallisation were merely dried at room temperature as a thin layer of about 2mm until a solids content of at least 78% by weight was achieved, normally it was achieved in 2-3 days.

25 v. The form of the zeolite which was produced was established using standard XRD method.

Specific description of the invention

25 Examples of processes of the invention will now be given to illustrate but not limit the invention.

Seed preparation:

30 A Sample of zeolite P was produced using the following procedure to be used as seed in the subsequent examples.

35 1420g of 2M sodium hydroxide solution and 445g of commercial sodium aluminate solution (concentration 20% Na₂O, 20% Al₂O₃) (Na₂O/Al₂O₃ = 1.64) were placed in a 5 litre baffled flask connected to a reflux condenser.

The resultant solution was stirred and heated to 90°C. 450g of commercial sodium silicate solution ((SiO₂ 28.3%/13.8% Na₂O w/w) SiO₂/Na₂O = 2:1) was diluted with 1100g of deionised water. The diluted silicate solution was heated to 75°C and added to the stirred aluminate caustic solution over 18 minutes. The resultant reaction mixture gel was allowed to react at 90°C with stirring for 5 hours. The product was filtered, washed and dried.

Example 1

40 Solutions A, B and C were prepared.

45 Solution A - 1000g of 2M sodium hydroxide solution.

50 Solution B - 490g of commercial sodium silicate solution as used in the seed preparation 300g of 4M sodium hydroxide solution 20g of P zeolite seed slurried in 30g deionised water.

55 Solution C - 588g of commercial sodium aluminate solution as used in the seed preparation. 300g of 4M sodium hydroxide solution.

60 Solution A was placed in a 5 litre round bottomed baffled flask with pitch blade turbine (500rpm) having a reflux condenser and heated to 90°C with vigorous stirring. Solution B and solution C were first preheated to 75°C and added simultaneously to solution A with solution B beginning 30 seconds before solution C, over 18 minutes and 37 minutes, respectively. The reaction gel mixture was allowed to react at 90°C with stirring for 5 hours. The product was filtered, washed and dried.

Example 2 (comparative)

65 The procedure of Example 1 was repeated with the omission of the P zeolite seed slurry.

Example 3

5 The procedure of Example 1 was repeated using a 2m³ vessel with suitable increased quantities of reactants. The vessel had a diameter of 1.13m and 2m height. The vessel had two stirrers. The bottom stirrer was a flat blade impeller and the middle stirrer was a pitch blade impeller (120 rpm with both stirrers on a single shaft).

Example 4

10 The procedure of Example 1 was repeated using a 20m³ vessel with a bottom flat bladed disc turbine stirrer (100 rpm) with the following increased quantities of reactants.

Solution A - 2,553 kg of 2M sodium hydroxide solution

15 Solution B - 3,724 kg of sodium silicate solution (conc. SiO₂ 14.7%. Na₂O 12.3%)

Solution C - 4,355 kg of sodium aluminate solution (conc. Na₂O 16.11%. Al₂O₃ 10.7%)

Solution D - 75 kg of P zeolite seed in 125 kg water.

20 The procedure of Example 1 was followed with solution D being added with solution B over 18 minutes.

Example 5

25 Solutions A, B and C were prepared.

Solution A - 648g of 2M sodium hydroxide solution

30 Solution B - 952g of commercial sodium silicate solution as used in the seed reaction

- 470g of 2M sodium hydroxide solution

- 20 g of P zeolite seed slurried in 30g deionised water

35 Solution C - 1139.5g of commercial sodium aluminate (20% Na₂O, 20% Al₂O₃)

- 805g of 2M NaOH solution

40 Solution A was placed in a 5 litre round bottomed baffled flask and heated to 90°C with stirring using the turbine of Example 1. Solution B and solution C were first preheated to 80°C and added, beginning simultaneously, to solution A over 20 minutes and 40 minutes respectively. The reaction gel was allowed to react at 90°C with stirring for 5 hours. The product was filtered, washed and dried.

Example 6 (Comparative)

45 The procedure of example 5 was repeated with the omission of the P zeolite seed slurry. The properties of these zeolite products are given in Table I.

Example 7

50 The procedure of Example 5 was repeated with the level of seed reduced by half (i.e. 10g of P zeolite, instead of 20g, was used in solution B).

Example 8

55 The procedure of Example 5 was repeated with 65g of the crystallised slurry obtained in Example 5 used as seed.

Example 9

The procedure of Example 5 was repeated with 30g of the washed filter cake of Example 5 used as seed.

5 Example 10

The procedure of Example 5 was repeated with 15g of the dried filter cake of Example 5 used as seed. Examples 8, 9 and 10 prove that the origin of the seed is of no importance.

10 Example 11

The procedure of Example 5 was repeated using a 20m³ vessel with a proportional increase in quantities of reactants.

15 Example 12

The procedure of Example 8 was repeated using a 20m³ vessel with a proportional increase in quantities of reactants.

20 Example 13

The procedure of Example 9 was repeated using a 20m³ vessel with a proportional increase in quantities of reactants.

25 Example 14

The procedure of Example 10 was repeated using a 20m³ vessel with a proportional increase in quantities of reactants.

30 Examples 4 and 11 to 14 prove that the process of the invention can be run at an industrial scale. That means that even with a very sensitive gel composition, most favorable for the production of zeolite A, it is nevertheless completely possible, with the process according to the present invention, to produce zeolite P at an industrial scale in a vessel of 20m³.

Another advantage of the present invention is therefore to enable the production of zeolite P using gel compositions and factory installations normally used for the production of zeolite A.

35 Example 15

The procedure of Example 5 was repeated with the reaction mixture (after reaching the crystallisation temperature of 90°C) allowed to react for:

40

- a- 30 minutes
- b- 60 minutes
- c- 180 minutes

45 Example 15 proves that crystallisation time is not a critical parameter if kept above 0.1 hour.

Example 16

50 The procedure of Example 5 was repeated with the temperature of the reactant components as well as the crystallisation temperature reduced to 60°C.

Zeolite P was produced together with a small amount of zeolite A proving that, for the most sensitive gel composition wherein the silicon:aluminum ratio is equal to 1, a temperature of about 60°C is actually the minimum for producing zeolite P. In order to have a pure P zeolite product, without any zeolite A, a crystallisation temperature of at least 80°C has been found necessary for such a gel composition and, in order to minimize the crystallisation time, the temperature 55 has to be preferably above 85°C.

Example 17

The procedure of example 11 was first repeated.

5 After the completion of crystallisation, filtration and washing of product, the resultant mother liquor and wash water solution was collected.

10 Part of the collected dilute mother liquor solution, which is essentially a dilute alkaline solution containing 8.7% Na₂O, was concentrated by evaporation up to 21% Na₂O and used to dissolve aluminium trihydrate. The concentrated sodium aluminate thus produced was diluted with some of the dilute mother liquor to give a sodium aluminate solution with the same concentration as that of solution C given in Example 5.

15 Part of the remaining dilute mother liquor was also used to dilute concentrated commercial sodium silicate solution to produce a sodium silicate solution having the same concentration as that of solution B given in Example 5.

The rest of the remaining dilute mother liquor played the same role as that of solution A given in Example 5.

20 The procedure of Example 11 was repeated using the solutions (A, B and C) prepared with the recycled mother liquor.

25 A product was obtained having the same characteristics as those of the product disclosed in Example 11 proving that it is completely possible to recycle the mother liquor.

Finally when comparing the oil absorption of the zeolites obtained according to the process of the present invention with the oil absorption of the comparative Examples or the oil absorption of the seed, it is immediately obvious that the oil absorption of a zeolite obtained according to the present invention is significantly higher.

20

25

30

35

40

45

50

55

Table I
Zeolite Properties

	Seed	Ex 1	Ex 2*	Ex 3	Ex 4	Ex 5	Ex 6*
$\text{SiO}_2/\text{Al}_2\text{O}_3$ (gel)	2.50	2.00	2.00	2.00	2.00	2	2
$\text{SiO}_2/\text{Al}_2\text{O}_3$ (product)	2.15	2.00	2.00	2.00	2.00	2	2
Yield % On Dry Basis	7.4	12.3	12.3	12.3	12.3	16.0	16.00
Zeolite	P	P	A + Hydroxy Sodalite (Trace)	P	P	P	A
APS (micron) d_{50}	0.90	0.94	5.0	0.90	0.90	0.95	7
CUR (secs)	8	4	>200	6	5	4	>200
CEBC (mg/g)	159	168	120	165	166	167	120
Oil Absorption (g oil/100g zeolite)	70	88	52	85	84	94	53

* comparative examples

These results demonstrate the process of the invention provides product zeolites having satisfactory properties. In the absence of the P-zeolite seed the product consisted of A zeolite as the sole or major product.

Table I (continued)

Zeolite Properties

	Ex 7	Ex 8	Ex 9	Ex 10	Ex 11	Ex 12	Ex 13
SiO ₂ /Al ₂ O ₃ (gel)	2	2	2	2	2	2	2
SiO ₂ /Al ₂ O ₃ (product)	2	2	2	2	2	2	2
Zeolite type	P	P	P	P	P	P	P
APS (micron) d ₅₀	0.90	0.95	0.88	0.85	0.96	0.90	0.93
CUR (secs)	5	6	7	8	6	5	7
CEBC (mg/g)	165	164	163	166	163	165	167
Oil Absorption (g oil/100g zeolite)	90	91	93	90	94	93	92

Table I (continued)

Zeolite Properties

	Ex 14	Ex 15a	Ex 15b	Ex 15c	Ex 17
SiO ₂ /Al ₂ O ₃ (gel)	2	2	2	2	2
SiO ₂ /Al ₂ O ₃ (product)	2	2	2	2	2
Zeolite Type	P	P	P	P	P
APS (micron) d ₅₀	0.94	1.5	1.1	0.95	0.95
CUR (secs)	6	7	5	6	7
CEBC (mg/g)	164	165	166	165	163
Oil Absorption (g oil/100g zeolite)	91	89	90	92	90

Claims

1. A process for the preparation of P zeolite having the oxide formula $M_n^2O \cdot Al_2O_3 \cdot (1.80-2.66)SiO_2 \cdot yH_2O$ wherein y is the water content, comprising the steps of:

i. a sodium aluminate solution having a temperature of at least 25°C is mixed with a sodium silicate solution at a temperature of at least 25°C in a stirred vessel in the presence of a slurry of P zeolite seed to form a gel having the composition,

Al₂O₃: (1.00-3.5) SiO₂: (1.2-7.5) Na₂O:
(25 to 450) (preferably 40 to 150) H₂O

5 ii. ageing the gel at a temperature above about 25°C with stirring to maintain solids in suspension for a period of at least about 0.1 hour, and

iii. separating the P zeolite product, washing and drying.

10 2. A process according to claim 1 for the preparation of P zeolite having the oxide formula $M_{2/n}O \cdot Al_2O_3 \cdot (1.90-2.10) SiO_2 \cdot yH_2O$ wherein y is the water content, comprising the steps of:

15 i. a sodium aluminate solution having a temperature of at least 60°C is mixed with a sodium silicate solution at a temperature of at least 60°C in a stirred vessel in the presence of a slurry of P zeolite seed to form a gel having the composition,

Al₂O₃: (1.80-2.2) SiO₂: (1.5-5) Na₂O: (40 to 150) H₂O

ii. ageing the gel at a temperature above about 60°C with stirring to maintain solids in suspension for a period of least about 0.1 hour, and

20 iii. separating the P zeolite product, washing and drying.

3. A process according to claim 2 wherein the resulting dilute mother liquor obtained after separating and washing the P zeolite product is collected:

25 i) a first part of it being concentrated and used to dissolve aluminium trihydrate to produce concentrated sodium aluminate,

30 ii) the obtained concentrated sodium aluminate being diluted with a second part of the dilute mother liquor giving a sodium aluminate solution,

iii) a third part of the dilute mother liquor being used to dilute a sodium silicate solution,

35 iv) the resulting sodium aluminate solution being then mixed with the sodium silicate solution in a stirred vessel containing the remaining part of the dilute mother liquor in the presence of a slurry of P zeolite seed to form a gel having the composition,

Al₂O₃: (1.80-2.2) SiO₂: (1.5-5) Na₂O: (40 to 150) H₂O

Patentansprüche

40 1. Verfahren zur Herstellung von P-Zeolith mit der Oxidformel

$M_{2/n}O \cdot Al_2O_3 \cdot (1.80-2.66) SiO_2 \cdot yH_2O$

45 worin y der Wassergehalt ist, mit den Stufen

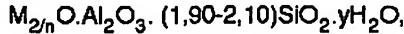
50 i. eine Natriumaluminatlösung mit einer Temperatur von wenigstens 25°C wird mit einer Natriumsilicatlösung bei einer Temperatur von wenigstens 25°C in einem Rührgefäß in Gegenwart einer Aufschämmung von P-Zeolith-Impfmaterial vermischt, um ein Gel zu bilden, das die Zusammensetzung aufweist,

Al₂O₃: (1,00-3,5) SiO₂: (1,2-7,5) Na₂O:
(25 bis 450) (vorzugsweise 40 bis 150) H₂O

55 ii. Altern des Gels bei einer Temperatur oberhalb von etwa 25°C unter Rühren, um Feststoffe in Suspension zu halten, für einen Zeitraum von wenigstens etwa 0,1 Stunde, und

iii. Abtrennen des P-Zeolith-Produkts, Waschen und Trocknen.

2. Verfahren nach Anspruch 1 zur Herstellung von P-Zeolith mit der Oxidform



5

worin y der Wassergehalt ist, mit den Stufen:

10 i. eine Natriumaluminatlösung mit einer Temperatur von wenigstens 60°C wird mit einer Natriumsilicatlösung bei einer Temperatur von wenigstens 60°C in einem Rührgefäß in Gegenwart einer Aufschämmung von P-Zeolith-Impfmaterial vermischt, um ein Gel zu bilden, das die Zusammensetzung aufweist
 $Al_2O_3: (1,80-2,2) SiO_2: (1,5-5,5) Na_2O: (40 \text{ bis } 150) H_2O$

15 ii. Altern des Gels bei einer Temperatur oberhalb etwa 60°C unter Röhren, um Feststoffe in Suspension zu halten, für einen Zeitraum von wenigstens etwa 0,1 Stunde, und

15 iii. Abtrennen des P-Zeolith-Produkts, Waschen und Trocknen.

20 3. Verfahren nach Anspruch 2, bei dem die erhaltene verdünnte Mutterlauge, die nach dem Abtrennen und Waschen des P-Zeolith-Produkts erhalten wird, gesammelt wird, wobei:

25 i) ein erster Teil davon konzentriert wird und dazu verwendet wird, Aluminiumtrihydrat aufzulösen, um konzentriertes Natriumaluminat herzustellen,

25 ii) das erhaltene konzentrierte Natriumaluminat mit einem zweiten Teil der verdünnten Mutterlauge verdünnt wird, was eine Natriumaluminatlösung ergibt,

30 iii) ein dritter Teil der verdünnten Mutterlauge dazu verwendet wird, eine Natriumsilicatlösung zu verdünnen,

30 iv) die erhaltene Natriumaluminatlösung danach mit der Natriumsilicatlösung in einem Rührgefäß vermischt wird, das den restlichen Teil der verdünnten Mutterlauge enthält, und zwar in Gegenwart einer Aufschämmung von P-Zeolith-Impfmaterial, um ein Gel zu bilden, das die Zusammensetzung aufweist
 $Al_2O_3: (1,80-2,2) SiO_2: (1,5-5,5) Na_2O: (40 \text{ bis } 150) H_2O$.

35 Revendications

1. Procédé de préparation de zéolite P ayant la formule d'oxyde $M_{2n}^2 \cdot Al_2O_3 \cdot (1,80 - 2,66) SiO_2 \cdot yH_2O$ dans laquelle y est la teneur en eau, comprenant les étapes consistant :

40 i. à mélanger une solution d'aluminate de sodium à une température d'eau moins 25°C à une solution de silicate de sodium à une température d'eau moins 25°C dans un récipient agité vigoureusement de manière adéquate en présence d'une quantité efficace d'une suspension de semence de zéolite P pour former un gel ayant la composition
 $Al_2O_3: (1,00-3,5) SiO_2: (1,2-7,5) Na_2O: (25 \text{ à } 450)$ (de préférence 40 à 150) H_2O ,

45 ii. à faire vieillir le gel à une température supérieure à environ 25°C sous agitation vigoureuse adéquate pour maintenir les solides en suspension pendant une période d'eau moins 0,1 heure, et
iii. à séparer la zéolite P produite, à la laver et à la sécher.

50 2. Procédé selon la revendication 1 pour la préparation de zéolite P ayant la formule d'oxyde
 $M_{2n}^2 \cdot Al_2O_3 \cdot (1,90 - 2,10) SiO_2 \cdot yH_2O$ dans laquelle y est la teneur en eau, le procédé comprenant les étapes consistant :

55 i. à mélanger une solution d'aluminate de sodium ayant une température d'eau moins 60°C à une solution de silicate de sodium à une température d'eau moins 60°C dans un récipient agité en présence d'une quantité efficace d'une suspension de semence de zéolite P pour former un gel de composition
 $Al_2O_3: (1,80-2,2) SiO_2: (1,5-5) Na_2O: (40 \text{ à } 150) H_2O$

55 ii. à faire vieillir le gel à une température supérieure à environ 60°C tout en l'agitant afin de maintenir les solides en suspension pendant une période d'eau moins 0,1 heure, et

iii. à séparer la zéolite P produite, à la laver et à la sécher.

5 3. Procédé selon la revendication 2, dans lequel la liqueur mère diluée obtenue après séparation et lavage de la zéolite P obtenue est recueillie :

10 i. une première partie de cette liqueur mère est concentrée et utilisée pour dissoudre de l'aluminium trihydraté afin de produire de l'aluminat de sodium concentré,

15 ii. l'aluminat de sodium concentré obtenu est dilué avec une deuxième partie de la liqueur mère diluée pour donner une solution d'aluminat de sodium,

iii. une troisième partie de la liqueur mère non diluée est utilisée pour diluer une solution de silicate de sodium,

iv. la solution d'aluminat de sodium obtenue est mélangée à la solution de silicate de sodium dans un récipient contenant la partie restante de la liqueur mère diluée en présence d'une suspension d'une semence de zéolite P pour former un gel ayant la composition

15 $\text{Al}_2\text{O}_3:(1,80-2,2) \text{ SiO}_2:(1,5-5) \text{ Na}_2\text{O}:(40 \text{ à } 150) \text{ H}_2\text{O}$.

20

25

30

35

40

45

50

55